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Protons + # Neutrons = Atomic mass #

(this is usually shown as an average for

all isotopes of an element) # Protons =



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Atomic Number = # Electrons = Protons  
(atoms of an element are electrically  
neutral so  $+$  =  $-$ ) # Neutrons = Atomic  
Mass - Atomic Number OR Atomic Mass  
- # of Protons

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Practice Worksheet Fill in the blanks in the following worksheet. Please keep in mind that the isotope represented by each space may NOT be the most common isotope or the one closest in atomic mass to the value on the periodic table. Atomic symbol Atomic number Protons Neutrons Electrons Atomic mass  
B 6

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arranged in specific areas of the atom  
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Atoms and Ions Worksheet Fill in the missing spaces in the charts below.

Element	Symbol	Atomic Number	Mass Number	Protons	Neutrons	Electrons
Carbon	C	6	14	6	8	6
Potassium	K	19	40	19	21	19
Tin	Sn	50	118	50	68	50
Zinc	Zn	30	66	30	36	30

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