

## Handbook Of Chemical Engineering Calculations Solution Manual

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Source: Otto Frank, "Shortcuts for Distillation Design," Chemical Engineering, March 14, 1977. In the present case, hand calculation and shortcut procedures are adequate because the benzene- toluene-xylene system is close to ideal. 2. Establish the equilibrium relationship among the constituents.

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Source: HANDBOOK OF CHEMICAL ENGINEERING CALCULATIONS. 1.2 SECTION ONE 1.1 MOLAR GAS CONSTANT Calculate the molar gas constant R in the following units: a. (atm)(cm<sup>3</sup>)/(g-mol)(K) b.

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Calculation Procedure 1. Assume a basis. Assume gas is at standard conditions, that is, 1 g-mol gas at 1 atm (101.3 kPa) pressure and 0 C (273 K, or 492 R), occupying a volume of 22.4 L. 2. Compute the gas constant. Apply suitable conversion factors and obtain the gas constant in various units. Use PV = RT; that is, R = PV/T. Thus, a.

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5. Bretzsnajder—Prediction of Transport and Other Physical Properties of Fluids, Pergamon. 6. Smith and Van Ness—Chemical Engineering Thermodynamics, McGraw-Hill. 7. Perry and Chilton—Chemical Engineers' Handbook, McGraw-Hill. 8. The Chemical Rubber Company—Handbook of Chemistry and Physics. 9.

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